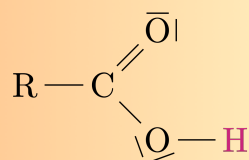
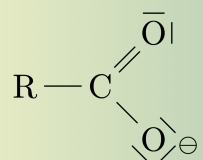
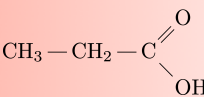
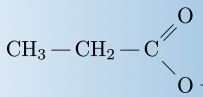
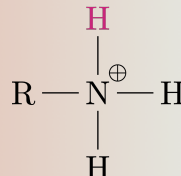
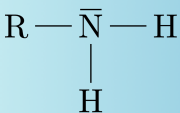
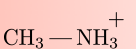
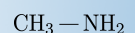


$$\text{pH} = -\log \left(\frac{[\text{H}_3\text{O}^+]}{c^\circ} \right)$$



$$[\text{H}_3\text{O}^+] = c^\circ \times 10^{-\text{pH}}$$

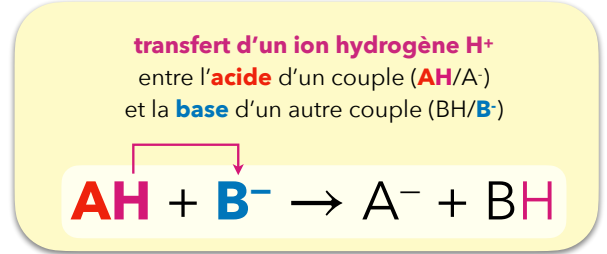
$$c^\circ = 1 \text{ mol} \cdot \text{L}^{-1}$$

Définition d'un acide selon Brønsted : espèce capable de céder un ion hydrogène H⁺	couple acide-base AH / A⁻	Définition d'une base selon Brønsted : espèce capable de capter un ion hydrogène H⁺
acide carboxylique 		ion carboxylate 
acide propanoïque 		ion propanoate 
ion ammonium 		amine 
ion méthylammonium 		méthylamine 

couples de l'**eau** :
 (H₃O⁺ / H₂O)
 (H₂O / HO⁻)
 couples de l'**acide carbonique** :
 (CO₂, H₂O / HCO₃⁻)
 (HCO₃⁻ / CO₃²⁻)

espèce **amphotère** — adj.
 un **ampholyte** — nom

Réaction acide-base :



Exemple :

Couples :
 (CH₃COOH / CH₃COO⁻)
 (H₂O / HO⁻)

